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CONJUGATE ACID-BASE PAIRS

In the reaction, Bronsted-Lowry Acid and Base, it was shown that after an acid has given up its proton, it is capable of getting back that proton and becoming a base. Conjugate base is what is left after an acid gives up a proton. How strong the acid, the weaker the conjugate base. The weaker the acid, the stronger the conjugate base.

Write the conjugate to the substances.

ACID	BASE	EQUATION
H_2SO_4	HSO_4^-	$H_2SO_4 \rightleftharpoons H^+ + HSO_4^-$
HNO_3	NO_3^-	$HNO_3 \rightleftharpoons H^+ + NO_3^-$
HCl	Cl^-	$HCl \rightleftharpoons H^+ + Cl^-$
$H_2PO_4^-$	HPO_4^{2-}	$H_2PO_4^- \rightleftharpoons H^+ + HPO_4^{2-}$
HPO_4^{2-}	PO_4^{3-}	$HPO_4^{2-} \rightleftharpoons H^+ + PO_4^{3-}$
H_2O	OH^-	$H_2O \rightleftharpoons H^+ + OH^-$
H_2CO_3	HCO_3^-	$H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$
HCO_3^-	CO_3^{2-}	$HCO_3^- \rightleftharpoons H^+ + CO_3^{2-}$
NH_4^+	NH_3	$NH_4^+ \rightleftharpoons H^+ + NH_3$
H_3O^+	H_2O	$H_3O^+ \rightleftharpoons H^+ + H_2O$

Write the conjugate bases: HSO_4^- or $H_2PO_4^-$

Write the conjugate bases: Cl^- or NO_3^-

Class: X-Science Acids, Bases and Salts

CBSE NCERT Solutions for Class 10 Science Chapter 2

Back of Chapter Questions

1. A solution turns red litmus blue; its pH is likely to be:

(A) 1
(B) 4
(C) 5
(D) 10

Solution: (D)

A solution which turns the colour of red litmus to blue must be basic in nature. A basic solution has a pH value of greater than 7. The only option with pH greater than 7 is 10 which will be the correct answer.

2. A solution reacts with crushed egg-shells to give a gas that turns lime-water milky. The solution contains:

(A) NaCl
(B) HCl
(C) LiCl
(D) KCl

Solution: (B)

Egg shells are made up of calcium carbonate, which on reaction with HCl liberates CO_2 gas. When CO_2 reacts with lime water ($Ca(OH)_2$), it forms calcium carbonate which gives the solution a milky appearance.

$$CaCO_3 + 2HCl \rightarrow CaCl_2 + H_2O + CO_2$$

$$Ca(OH)_2 + CO_2 \rightarrow CaCO_3 + H_2O$$

3. 10 ml of a solution of NaOH is found to be completely neutralized by 8 mL of a given solution of HCl. If we take 20 ml of the same solution of NaOH, the amount of HCl solution (the same solution as before) required to neutralize it will be:

(A) 4 ml
(B) 8 ml
(C) 12 ml
(D) 16 ml

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Quiz & Worksheet - Bronsted-Lowry Acids

1. For the following reaction between NH_3 and CH_3COOH , which one is the Bronsted-Lowry acid?

$$NH_3 + CH_3COOH \rightleftharpoons NH_4^+ + CH_3COO^-$$

NH_3
 CH_3COOH
 NH_4^+
 CH_3COO^-

2. Which of the following is a correct conjugate acid-base pair?

$$HF + NH_3 \rightleftharpoons NH_4^+ + F^-$$

HF/NH_3
 HF/NH_4^+
 NH_3/F^-
 HF/F^-

3. Which of the substances in the following chemical reaction is the conjugate base?

$$H_3O^+ + NH_3 \rightleftharpoons NH_4^+ + H_2O$$

H_3O^+
 NH_4^+
 NH_3
 H_2O

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